

The Crystal Structure of the Basic Mercury(II) Perchlorate $\text{Hg}_2\text{OOHClO}_4$

GEORG JOHANSSON and ERNST HANSEN

*Department of Inorganic Chemistry, Royal Institute of Technology,
S-100 44 Stockholm 70, Sweden*

The crystal structure of $\text{Hg}_2\text{OOHClO}_4$ has been determined from three-dimensional X-ray data using visually estimated intensities from Weissenberg photographs. The unit cell is monoclinic with $a = 11.078$ Å, $b = 14.269$ Å, $c = 7.224$ Å, and $\beta = 100.5^\circ$. The space group is No. 14: $P2_1/c$. Each Hg atom forms two short Hg-O bonds with lengths of 2.0 Å and with the O-Hg-O angle close to 180° . Other Hg-O contacts are longer than about 2.7 Å. The structure can be described as built up from infinite one-dimensional complexes in which the Hg atoms are joined by the covalently bonded oxygens. The characteristic OHg_3 groups, previously found in several other structures, can be distinguished in these complexes.

In connection with an X-ray investigation of the structures of the polynuclear hydrolysis complexes of mercury(II) in aqueous perchlorate solutions,¹ it was found of interest to determine the structures of the basic salts that could be crystallized from such solutions. Of particular interest is the coordination of the mercury atom in the crystals and the type of bridging between the metal atoms in the polynuclear complexes. The structure of the basic perchlorate $\text{Hg}_5\text{O}_2(\text{OH})_2(\text{ClO}_4)_4(\text{H}_2\text{O})_x$, which is the least basic of those obtained, has been described in a previous paper.² The crystal structure of the most basic of the salts, $\text{Hg}_2\text{OOHClO}_4$, is given in the present paper, and an intermediate phase, $\text{Hg}_7\text{O}_4(\text{OH})_2(\text{ClO}_4)_4$, will be discussed in a following paper.

EXPERIMENTAL

Preparation of crystals. Solutions obtained by dissolving an excess of HgO in HClO_4 were used for the preparation of the crystals. Needlelike or rod-shaped crystals suitable for X-ray work crystallize spontaneously from a solution which is 3.1 M in Hg^{2+} and 4.8 M in ClO_4^- . The crystals are stable outside the mother liquor but decompose to HgO when washed with water.

Analysis. The crystals were separated from the mother liquor by centrifuging and were dried between filter papers. Dilute perchloric acid was used to dissolve the crystals before analysis.

Mercury was determined by reducing with H_3PO_3 and precipitating and weighing as Hg_2Cl_2 . Perchlorate was determined by passing a portion of a solution through a cation exchanger and titrating the eluate with NaOH , correcting for the amount of HClO_4 used to dissolve the crystals. The density was calculated from the apparent loss of weight in benzene.

The results of the analysis are compared in Table 1 with values calculated for the stoichiometric composition $4\text{HgO}\cdot\text{Cl}_2\text{O}_7\cdot\text{H}_2\text{O}$. A basic mercury(II) perchlorate, probably identical with the one described here, has been prepared previously by Hayek and Schnell.³ They assumed the composition to be $4\text{HgO}\cdot\text{Cl}_2\text{O}_7$, and the results of their analysis are given in Table 1 for comparison.

Table 1. Analytical data.

	Found	Calculated for $4\text{HgO}\cdot\text{Cl}_2\text{O}_7\cdot\text{H}_2\text{O}$	Found by Hayek and Schnell ³
% HgO	79.0	81.3	81.3
% Cl_2O_7	18.2	17.1	18.1
% H_2O	2.8 (diff.)	1.7	
Density (g cm^{-3})	6.1 ₃	6.28	

For a direct estimation of the number of non-perchlorate oxygens a Karl Fischer titration⁴ was carried out. It can be expected that basic oxygens (O^{2-} and OH^-) will react as water in this titration and the result will give the total number of oxygens not belonging to the perchlorate groups. Eight different titrations gave an average value of 0.9₄ such oxygens per mercury atom. It will be shown in the following, that the crystal structure determination indicates that the correct formula is $\text{Hg}_2\text{OOHClO}_4$. This is consistent with the number of oxygens estimated from the Karl Fischer titration and corresponds to the stoichiometric composition $4\text{HgO}\cdot\text{Cl}_2\text{O}_7\cdot\text{H}_2\text{O}$.

Unit cell and space group. Weissenberg photographs taken along two of the axes of the unit cell showed the symmetry to be monoclinic. More accurate values for the unit cell dimensions than those obtained from the Weissenberg photographs were determined by a least squares refinement using the line positions on a powder photograph taken in a Guinier camera. $\text{CuK}\alpha$ radiation ($\lambda = 1.5405 \text{ \AA}$) was used with KCl ($a = 6.2929 \text{ \AA}$) as internal standard. Observed and calculated line positions for the first part of the powder photograph are given in Table 2. The unit cell dimensions were found to be:

$$a = 11.07_8 \text{ \AA}, b = 14.26_6 \text{ \AA}, c = 7.22_4 \text{ \AA}, \beta = 100.5^\circ$$

The calculated density, assuming four formula weights $4\text{HgO}\cdot\text{Cl}_2\text{O}_7\cdot\text{H}_2\text{O}$ in the unit cell, is 6.28. The observed value is 6.1₃ (Table 1). Systematically absent reflections are $h0l$ for $l = 2n + 1$ and $0k0$ for $k = 2n + 1$. This is consistent with the centrosymmetric space group No. 14, $P2_1/c$.

Intensity data. Intensity data were collected in a Weissenberg camera with the use of $\text{MoK}\alpha$ radiation. Crystals with as far as possible cylindrical cross sections, with the largest dimension not exceeding 0.10 mm, were selected. Photographs were taken around the c axis ($hk0 - hk8$) and the b axis ($h0l - h4l$). Intensities were determined by visual comparison with an intensity scale prepared by timed exposures of one of the reflections of the same crystal. Intensities for about 2350 independent reflections were determined and were used for the structure determination. Lorentz and polarization factors were applied in the usual way and all intensities were recalculated to a common scale with the use of reflections appearing on more than one photograph.

Computer programs. The following programs were used to carry out the calculations: DRF. Lp corrections, structure factor calculations and Fourier summations. Originally written by A. Zalkin, Berkeley, USA. Modified by R. Liminga and J. O. Lundgren, Uppsala, Sweden.

Table 2. Calculated and observed $\sin^2\theta$ values for the first lines of a Guinier powder photograph of $\text{Hg}_2\text{OOHClO}_4$.

hkl	$(\sin^2\theta)_{\text{calc}}$	$(\sin^2\theta)_{\text{obs}}$	I_{obs}
110	0.00792	0.00789	vw
020	0.01166	0.01165	vw
011	0.01467	0.01465	m
111	0.01680	0.01677	s
200	0.02000	0.01998	w
111	0.02255	0.02251	m
210	0.02292	0.02290	vw
021	0.02341	0.02342	vw
121	0.03129	0.03126	w
220	0.03166	0.03163	w
221	0.03767	0.03766	m
031	0.03799	0.03796	vw
300	0.04501	0.04498	s
310	0.04792	0.04792	vw
221	0.04917)	0.04910	vw
112	0.04920)		
012	0.04995	0.04999	m
140	0.05163	0.05163	s
231	0.05224	0.05221	vw
202	0.05554	0.05555	vw
320	0.05666	0.05661	w
102	0.05778	0.05780	s
022	0.05869	0.05873	s
141	0.06051	0.06053	w
240	0.06663	0.06663	w
222	0.06719	0.06723	w
311	0.06830	0.06835	w
330	0.07123	0.07122	vw
241	0.07264	0.07262	vw
331	0.07437	0.07444	w
150	0.07785)	0.07781	vs
312	0.07771)		
400	0.08001	0.08013	vw
232	0.08177	0.08179	m
411	0.08319	0.08299	vw
132	0.08401)	0.08416	s
241	0.08414)		
322	0.08645	0.08646	w
222	0.09019	0.09013	vw
420	0.09167)	0.09170	vw
421	0.09193)		
250	0.09286	0.09279	vw
042	0.09366	0.09362	vw
341	0.09477	0.09482	vw
251	0.09887	0.09891	w
232	0.10476)	0.10489	vw
060	0.10491)		
113	0.10512)		

LALS. Full matrix least squares refinement. Originally written by P. K. Gantzel, R. A. Sparks and K. N. Trueblood, Los Angeles, USA. Modified by Å. Zalkin, Berkeley, USA, and by J.-O. Lundgren, R. Liminga and C. I. Brändén, Uppsala, Sweden.

DISTAN. Calculations of interatomic distances and bond angles. Written by A. Zalkin, Berkeley, USA, and modified by A. G. Nord and B. G. Brandt, Stockholm, Sweden.

STRUCTURE DETERMINATION

The three-dimensional Patterson function showed that the 16 Hg atoms in the unit cell must occupy the general four-fold positions. Possible parameter values were deduced from the Harker sections. A systematic comparison of calculated Hg–Hg vectors with peaks in the three-dimensional Patterson function, sufficiently high to be identified as possible Hg–Hg vectors, led to an apparently unique set of parameter values for the four four-fold Hg positions in the unit cell.

A least squares refinement of the derived positions using all observed reflections gave an R factor of 0.20 and led to reasonable temperature factors for the four independent Hg atoms. On the basis of the derived parameter values a three-dimensional difference map was calculated from which the chlorine atoms and most of the oxygen atoms could be located. A new least squares refinement followed by a three-dimensional difference map led to the positions of the remaining oxygen atoms.

With all atoms included, a new least squares refinement using all observed reflections and with individual isotropic temperature factors for all atoms lowered the R factor to 0.16. The scattering factors used were those given by Cromer and Waber⁵ for the neutral atoms. Corrections were made for the real part of the anomalous dispersion according to Cromer.⁶ The weighting scheme was that suggested by Hughes.⁷ A further least squares refinement

Table 3. Final parameters for $\text{Hg}_2\text{OOHClO}_4$. All atoms are in position 4(e) in space group No. 14; $P2_1/c$. Standard deviations are given within brackets.

Atom	x	y	z	B
Hg ₁	0.3207 (2)	0.2484 (1)	0.9784 (2)	1.1 (0)
Hg ₂	0.2649 (2)	0.0487 (1)	0.2020 (2)	1.2 (0)
Hg ₃	0.6744 (2)	0.1842 (1)	0.9093 (2)	1.5 (0)
Hg ₄	0.9654 (2)	0.1114 (1)	0.8376 (2)	1.5 (0)
Cl ₁	0.9604 (10)	0.1384 (8)	0.3462 (15)	1.7 (2)
O ₁	0.0179 (54)	0.0951 (44)	0.2192 (83)	4.5 (14)
O ₂	0.8802 (44)	0.1973 (38)	0.2621 (68)	2.7 (8)
O ₃	0.8921 (47)	0.0891 (40)	0.4314 (73)	4.1 (12)
O ₄	0.0495 (49)	0.1865 (40)	0.4756 (73)	3.1 (10)
Cl ₂	0.3999 (11)	0.4777 (9)	0.2354 (15)	1.8 (2)
O ₅	0.3056 (36)	0.4359 (30)	0.1011 (53)	1.6 (7)
O ₆	0.5195 (63)	0.4575 (53)	0.2045 (97)	4.5 (15)
O ₇	0.6071 (56)	0.0779 (50)	0.2629 (84)	4.5 (15)
O ₈	0.4032 (74)	0.4336 (63)	0.4399 (114)	5.8 (19)
O ₉	0.7805 (27)	0.4141 (22)	0.3069 (39)	1.9 (6)
O ₁₀	0.3341 (30)	0.1834 (23)	0.2254 (43)	1.7 (6)
O ₁₁	0.5754 (37)	0.2685 (30)	0.0350 (54)	3.1 (10)
O ₁₂	0.1422 (36)	0.3822 (29)	0.3815 (54)	2.9 (9)

Table 4. Observed and calculated structure factors.

0 k 0	0	5 k 0	14	74 - 74	13	180 -184	14	61 - 59	3	61 - 76	16	99 - 95			
0	1800	0	254 -255	15	< 52 10	14	85 80	15	< 41 - 23	4	108 140	17	< 48 22		
2	125 -119	1	62 93	16	150 102	15	51 50	16	109 -125	5	211 278	18	< 50 - 29		
4	98 31	2	199 241	17	< 55 38	16	< 42 9	17	104 108	6	76 75	19	149 174		
6	288 -259	3	52 - 34	18	120 -115	17	121 142	18	52 46	7	50 40		9 k		
8	75 - 31	4	229 -247	19	10 k 0	18	127 134	19	< 48 - 26	8	< 34 - 7	1	119 -144		
10	167 -131	5	63 - 75	0	100 - 95	20	< 49 41	21	114 142	10	36 39	3	127 -154		
12	< 40 - 18	6	< 32 71	1	112 107	21	57 56		-13 k 1	11	37 36	4	< 37 36		
14	378 -433	7	442 -504	2	57 - 50				1	90 - 82	13	< 40 - 29	5	167 198	
16	265 284	8	127 133	3	189 203	1 k 1	1	232 -253	2	108 -112	14	106 -133	7	< 38 21	
18	50 - 11	9	165 174	4	88 95	1	252 200	3	455 -352	5	< 44 21	8	< 39 1	8	62 - 67
20	136 157	10	170 180	5	98 - 92	2	152 85	4	278 207	15	151 -181	9	< 47 28	9	62 - 67
		11	55 42	6	< 43 8	3	112 85	5	< 27 - 30	17	< 54 - 28	11	142 135	10	< 39 7
0	< 12 9	12	92 - 97	7	183 157	4	150 -105	5	< 27 - 30	17	< 47 28	12	< 43 - 28	11	142 135
2	60 - 65	13	127 125	8	< 45 19	5	295 -257	6	66 71	18	54 - 28	11	142 135	12	< 43 - 28
2	191 139	14	128 145	9	< 143 -117	6	197 -175	7	63 46	19	89 -105	12	< 43 - 28	12	< 43 - 28
3	< 14 - 10	15	< 47 41	10	< 47 3	7	213 179	8	130 107	19	59 - 28	13	< 43 - 28	13	< 43 - 28
4	473 499	16	167 -161	11	172 136	8	149 -136	9	258 236	1	41 40	14	< 46 - 57	14	< 46 - 57
5	625 597	17	< 50 53	12	110 88	9	366 345	10	63 34	1	116 -109	15	< 47 51	15	< 47 51
6	87 - 74	18	92 83	13	< 51 - 53	10	74 - 42	11	293 264	2	338 -330	16	< 49 - 61	16	< 49 - 61
7	130 109	19	< 54 18	14	< 52 29	11	< 33 12	12	175 -167	3	44 30	17	< 50 - 29	17	< 50 - 29
8	179 195	20	79 - 85	15	119 - 97	12	238 -227	13	89 - 96	4	68 - 64	18	< 52 - 12	18	< 52 - 12
9	422 -416	21	81 104	16	< 55 - 10	13	163 152	14	< 39 - 17	5	201 -195	19	129 -142	19	129 -142
10	167 190	22	< 60 - 51	17	< 113 - 78	14	51 - 74	15	63 43	6	< 31 19				
11	227 239	23	109 -132	0	11 k 0	15	< 40 - 14	15	< 42 - 13	7	121 -118	1	171 k 1		
12	139 -160	24	84 57	0	241 -215	16	91 - 89	17	111 112	8	229 218	1	171 k 1		
13	49 57	0	184 209	1	54 38	17	101 - 91	18	100 - 88	9	122 117	2	76 - 64		
14	< 42 26	1	157 190	2	221 181	18	< 45 40	19	85 50	10	121 112	3	61 - 54		
15	127 134	2	348 -422	3	116 100	19	< 47 13	20	75 54	11	211 215	4	< 35 22		
16	77 - 77	3	165 208	4	< 44 28	20	75 54	21	148 -161	1	102 -150	12	< 38 13		
17	72 50	4	< 32 6	5	< 44 41	21	148 -161	2	131 181	14	133 130	6	244 -207		
18	126 135	5	87 97	6	< 45 15	1	-1 k 1	3	149 174	15	48 - 52	7	< 37 - 35		
19	132 135	6	< 7 8	7	< 45 - 80	1	310	4	212 -257	16	43 55	8	180 -139		
20	< 84 - 14	7	208 -207	8	129 -106	2	< 14 5	5	101 -112	17	86 90	9	< 39 13		
21	124 -138	8	111 108	9	< 48 49	3	< 79 51	6	194 -218	18	< 48 - 4	10	145 126		
22	< 57 - 50	9	125 123	10	109 - 74	4	421 -385	7	174 -127	19	79 - 69	11	< 41 21		
23	< 59 12	10	90 94	11	< 52 16	5	174 -127	8	226 -225	19	79 - 69	11	< 41 21		
24	< 61 - 27	11	< 41 14	12	117 - 98	6	261 225	9	45 34	1	< 30 38	13	115 -101		
25	133 -161	12	209 203	13	< 53 - 30	7	143 -107	10	< 36 - 66	2	100 147	14	110 104		
0	218 -236	13	< 44 15	14	126 113	8	154 -130	11	120 122	2	162 112	15	< 47 - 31		
2	101 102	14	155 -175	15	< 55 - 49	9	68 - 48	12	243 267	3	102 127	16	< 48 19		
3	49 56	15	125 -147	16	102 - 94	10	258 -224	13	162 146	4	91 -123	17	89 76		
4	147 -449	16	< 55 23	17	< 58 - 50	11	101 -101	14	< 40 15	5	< 51 - 5	18	< 51 7		
5	337 302	17	81 - 53	18	84 73	12	121 114	15	< 42 15	6	< 33 - 44	19	< 53 4		
6	< 27 26	20	< 57 - 20	0	12 k 0	13	217 -204	16	< 43 0	7	48 - 39	20	36 87		
7	357 -328	21	82 84	1	71 55	14	< 37 11	17	48 - 74	8	49 - 50	1	10 k 1		
8	322 312	0	7 k 0	2	78 - 71	15	116 110	18	81 103	9	175 -163	1	37 - 50		
9	314 289	0	9 k 0	3	64 - 55	16	110 108	19	103 104	10	189 -225	2	105 114		
10	317 332	1	114 106	4	48 131	19	134 134	1	327 328	12	132 156	3	110 129		
11	87 - 82	2	80 102	5	< 45 - 3	20	107 -109	2	255 240	14	57 - 39	5	98 55		
12	66 - 72	3	329 309	6	< 48 38	1	2 k 1	3	53 - 35	15	< 45 31	6	109 133		
13	63 56	4	41 45	7	200 -171	2	64 113	4	46 46	16	< 46 - 28	7	< 40 15		
14	119 119	5	157 -160	9	90 72	3	72 - 76	6	358 298	18	74 67	9	97 -100		
15	58 - 48	6	56 - 63	10	95 - 65	4	437 495	8	203 198	19	91 71	10	111 -127		
16	143 -148	7	260 240	11	< 52 44	5	58 - 51	9	48 48	1	80 91	12	50 37		
17	< 47 - 16	8	< 47 - 16	12	< 53 50	6	< 27 - 4	9	48 48	1	80 91	12	50 37		
18	79 79	9	134 -127	13	77 70	7	< 28 28	10	129 -115	1	339 274	13	83 72		
19	82 - 69	10	90 - 88	13 k 0	116 0	8	< 28 - 13	12	85 79	3	194 -172	14	80 - 79		
20	56 -124	12	90 - 92	0	116 -105	9	< 30 25	12	85 79	4	130 128				
21	135 135	12	90 - 92	0	116 -105	9	< 30 25	12	85 79	4	130 128				
22	93 -106	13	79 - 80	2	114 - 78	11	< 33 25	14	163 -154	5	6 160 156	2	144 128		
0	610 851	16	< 47 21	3	< 49 20	12	104 - 98	15	94 - 92	7	33 53	3	161 -154		
1	185 244	17	125 -115	4	87 77	13	< 35 29	16	96 109	8	< 34 33	4	100 - 91		
2	274 -321	19	< 56 82	5	163 -114	14	89 - 87	17	48 19	9	169 165	5	< 81 69		
3	217 246	18	< 54 46	7	< 51 60	16	208 -211	18	< 40 22	10	100 - 95	6	175 - 70		
4	64 33	20	< 58 45	8	< 52 7	17	< 44 6	19	102 90	10	100 - 95	6	175 - 70		
5	241 217	21	< 59 54	10	< 53 0	19	< 47 25	2	38 48	12	115 115	8	85 - 79		
6	128 -108	22	< 61 16	11	< 55 4	20	124 124	3	32 - 35	15	< 44 - 36	10	45 - 48		
7	59 - 61	23	138 124	12	119 92			4	191 251	16	120 -133	12	164 156		
8	62 64	0	8 k 0	0	14 k 0	1	< 30 - 27	5	< 28 27	16	120 -133	12	164 156		
9	145 129	0	9 k 0	1	205 194	2	343 -303	6	145 -166	1	107 -110	14	53 - 44		
10	56 - 38	1	< 35 40	1	51 - 4	3	275 221	7	99 106	1	99 106	1	99 106		
11	45 - 20	2	294 309	2	71 54	4	261 -182	8	86 89	3	51 67	11	k 1		
12	103 89	3	< 46 10	3	72 54	5	543 -405	9	49 48	4	< 35 23	1	39 - 5		
13	109 - 95	4	51 - 54	4	51 - 12	6	201 -143	10	186 204	5	53 70	2	107 -112		
14	316 -337	4	51 - 54	5	89 59	7	163 119	12	< 38 5	6	197 -215	3	116 118		
15	214 -198	5	87 - 91	6	< 53 17	8	139 -109	13	153 141	7	44 42	4	< 40 - 4		

BASIC MERCURY(II) PERCHLORATE

Table 4. Continued.

16	113	-109	3	196	190	10	<	90	-76	6	<	97	36	7	71	-60	12	127	-153	14	154	-81						
1	12 k 1	0	4	85	5	11	<	94	-105	7	156	130	8	146	138	13	164	194	15	<	77	12						
2	7	-82	5	244	230	12	<	98	-30	8	<	101	-80	9	198	-196	14	<	71	-33	16	138	-187					
3	3	-61	6	340	282	13	114	-121	0	9	>	102	26	10	94	84	15	<	74	23								
4	3	-91	8	240	-360					10	<	106	-71	11	127	-139	16	<	-270	16	<	77	12					
5	5	116	9	256	247	0	5	2		11	127	-139	12	13	146	144	17	127	-111	1	8 k 3	0						
6	7	<	10	82	-42	1	137	-121	0	12	-9	2	13	146	144	18	68	-69	1	242	-241	3	88	-97				
7	6	<	11	86	-83	2	306	359	0	13	85	75	16	14	68	-69	1	242	-241	3	88	-97	4	114	134			
8	6	<	12	169	-166	3	320	287	2	14	85	75	16	14	68	-69	1	242	-241	3	88	-97	5	62	51			
9	9	75	4	225	-192	4	225	-192	2	15	46	35	17	17	77	-81	4	251	189	6	109	109	6	109	109			
10	10	<	5	346	-284	5	346	-284	4	16	88	36	18	90	-108	5	296	238	7	>	65	-12	7	104	109			
11	10	<	6	185	-152	6	122	-104	5	17	88	36	18	90	-108	6	96	-75	8	80	-63	10	<	71	0			
12	11	101	7	1	36	24	7	144	-96	5	110	-110	19	93	66	7	75	60	9	54	-36	11	<	73	81			
13	11	101	8	210	-182	8	7	79	27	6	162	-181	20	93	66	8	80	-63	10	54	-36	11	<	73	81			
14	11	101	9	213	245	9	110	-115	7	155	153	1	193	-248	8	90	-63	10	54	-36	11	<	73	81				
15	11	101	10	5	21	-17	10	216	177	8	141	130	2	91	-121	9	54	-36	11	54	-36	11	<	73	81			
16	12	47	4	5	202	-151	11	105	210	9	149	-154	3	168	-174	10	153	126	12	321	-272	5	62	51				
17	4	137	115	7	181	142	12	298	-276	10	125	-141	4	60	-45	11	323	256	13	321	-272	5	62	51				
18	5	42	40	8	7	-21	13	6 k 2	12	11	<	105	83	5	448	387	12	17	51	1	-8 k 3	0	1	112	-128			
19	6	-12 k 1	9	9	76	-46	14	164	185	6	260	-221	13	6	260	-221	13	6	260	-221	13	6	260	-221	13	6	260	-221
20	7	150	204	10	204	-145	15	182	143	7	45	36	14	69	-32	3	321	-272	5	62	51	4	140	154				
21	8	78	66	11	202	-177	16	170	-188	8	10	2	10	2	10	2	10	2	10	2	10	2	10	2	10	2		
22	9	101	80	12	210	-190	17	4	271	342	0	<	40	-38	10	123	105	5	5	5	5	5	5	5	5			
23	10	116	140	13	146	-138	18	5	78	69	1	118	-113	12	89	72	11	103	128	7	91	-153	6	108	141			
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25	12	47	15	15	103	-37	20	7	84	56	3	56	56	14	151	135	13	94	4	4	4	4	4	4	4			
26	13	69	58	17	116	-226	21	8	93	114	4	93	114	14	151	135	14	94	4	4	4	4	4	4	4			
27	13	69	58	18	117	-219	22	9	90	53	5	100	-70	16	104	84	15	152	181	11	92	86	11	92	86			
28	14	87	71	19	149	115	23	10	206	-247	6	142	149	17	76	-58	16	85	78	1	9 k 3	0	1	9 k 3	0			
29	15	87	71	20	149	115	24	11	115	95	7	142	149	18	80	41	17	86	66	2	86	66	2	86	66			
30	16	47	45	21	168	161	25	12	432	453	8	106	106	19	166	-169	18	108	-248	3	64	28	3	64	28			
31	17	44	36	22	250	264	26	13	267	283	9	109	-113	20	176	-176	19	108	-248	3	64	28	4	108	-131			
32	18	44	36	23	250	264	27	14	267	283	10	109	-113	21	176	-176	20	108	-248	3	64	28	5	108	-131			
33	19	44	36	24	250	264	28	15	267	283	11	109	-113	22	176	-176	21	108	-248	3	64	28	6	108	-131			
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40	26	44	36	31	250	264	35	22	267	283	18	109	-113	29	176	-176	28	108	-248	3	64	28	13	108	-131			
41	27	44	36	32	250	264	36	23	267	283	19	109	-113	30	176	-176	29	108	-248	3	64	28	14	108	-131			
42	28	44	36	33	250	264	37	24	267	283	20	109	-113	31	176	-176	30	108	-248	3	64	28	15	108	-131			
43	29	44	36	34	250	264	38	25	267	283	21	109	-113	32	176	-176	31	108	-248	3	64	28	16	108	-131			
44	30	44	36	35	250	264	39	26	267	283	22	109	-113	33	176	-176	32	108	-248	3	64	28	17	108	-131			
45	31	44	36	36	250	264	40	27	267	283	23	109	-113	34	176	-176	33	108	-248	3	64	28	18	108	-131			
46	32	44	36	37	250	264	41	28	267	283	24	109	-113	35	176	-176	34	108	-248	3	64	28	19	108	-131			
47	33	44	36	38	250	264	42	29	267	283	25	109	-113	36	176	-176	35	108	-248	3	64	28	20	108	-131			
48	34	44	36	39	250	264	43	30	267	283	26	109	-113	37	176	-176	36	108	-248	3	64	28	21	108	-131			
49	35	44	36	40	250	264	44	31	267	283	27	109	-113	38	176	-176	37	108	-248	3	64	28	22	108	-131			
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52	38	44	36	43	250	264	47	34	267	283	30	109	-113	41	176	-176	40	108	-248	3	64	28	25	108	-131			
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56	42	44																										

Table 4. Continued.

0	56	56	12	151	158	0	< 51	13	8	< 105	78	7	70	-67	13	116	112	10	77	75	
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2	58	-45	14	< 69	-16	2	97	-103	11	< 76	-45	10	76	-61	16	177	160				
3	119	-132	15	135	-152	3	286	-343	12	< 78	-36	11	235	-226		4 k 5		2	-7 k 5	48	
4	350	372	16	< 75	6	4	79	-75	13	123	135	12	92	-84	1	205	213	2	< 35	-22	
5	297	274	17	< 70	94	5	< 55	16	0	-9 k 4		13	< 47	-16	2	168	-175	3	65	48	
6	213	-199	18	< 81	-92	6	6	103	-115	0	86	14	75	57	4	< 34	19	4	5	193	
7	350	323	19	94	-110	7	< 59	20	1	86	90	15	< 52	-49	4	< 35	40	5	193	158	
8	129	111				8	193	210	2	143	-170	17	< 56	25	5	93	-110	6	56	38	
9	< 51	35		-3 k 4		9	< 63	-40	3	57	54	18	< 50	49	6	104	-110	7	54	-45	
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14	< 67	-35	4	< 39	1	14	< 75	-25	8	231	235	1	38	89	11	79	-81	12	117	96	
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16	< 73	15	6	138	112	16	< 80	12	10	94	-98	5	35	-45	13	115	-125	14	54	65	
17	< 76	56	7	< 49	28	17	143	155	11	105	83	4	44	-41	14	51	62	15	< 55	42	
18	88	-83	8	< 52	-29	18	-6 k 4		12	162	170	5	45	55	15	62	-55	16	< 57	75	
19	111	-109	9	10	493	-433	0	-16 k 4	-172	10 k 4		6	189	-186	16	99	84	17	< 59	-12	
0	43	27	11	77	-53	1	128	131	0	51	40	7	< 34	-7	17	98	-50	18	107	115	
1	510	-502	12	95	77	2	123	-115	1	99	-101	8	< 35	-23	18	68	-50	19	151	144	
2	108	130	14	< 67	-31	4	373	333	3	< 68	45	10	< 43	15	1	-4 k 5		2	45	19	
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4	148	151	16	< 75	5	6	231	-193	5	< 70	-35	12	111	-115	3	50	46	5	< 47	12	
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7	75	71	19	< 82	-26	9	104	-82	8	181	208	15	< 32	-30	5	283	225	8	< 51	-43	
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9	128	-116				11	92	73	0	-10 k 4		17	89	89	7	116	-105	10	52	38	
10	201	-191	0	4 k 4		12	155	137	1	193	220	1	-1 k 5		8	116	-105	11	52	38	
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0	510	-439	12	< 65	-67	4	< 58	-47	13	145	170	13	96	-88	1	5 k 5		7	50	39	
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2	133	-161	14	< 71	-42	6	109	-122	15	90	93	15	< 51	-1	2	< 36	-20	11	< 49	-47	
3	216	-232	15	133	162	7	184	-224	0	11 k 4		16	< 56	-38	3	171	-187	12	99	-107	
4	121	-114	16	141	114	8	184	-224	1	54	66	17	< 50	-16	4	147	-162	13	< 53	-5	
5	217	190	0	178	-184	9	< 66	21	2	115	90	18	< 50	-16	5	< 39	-43	14	< 55	-26	
6	421	389	1	311	290	10	< 68	21	3	121	-140	19	< 60	-28	6	< 41	-33	15	< 57	-47	
7	117	89	2	34	0	11	< 70	2	4	121	-140	20	117	-124	7	< 42	23	16	< 59	0	
8	391	-302	2	34	0	12	< 72	32	5	< 74	43	21	2 k 5		8	< 42	23	16	< 59	0	
9	83	70	3	39	21	13	124	-143	6	< 76	29	1	61	-67	9	79	90	18	99	-97	
10	365	291	4	39	21	14	103	-83	7	< 76	29	2	61	-67	10	47	-48	19	9 k 5		
11	184	-154	5	324	267	0	73	67	8	< 77	-34	2	3	256	-276	11	148	171	1	< 48	-38
12	103	82	6	340	276	0	73	67	9	< 79	6	3	256	-276	11	148	171	1	< 48	-38	
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15	85	-67	9	161	121	3	63	-61	1	64	-61	14	164	163	14	< 54	52	3	< 48	-38	
16	< 72	-35	10	118	84	4	128	117	2	94	105	7	84	65	15	< 56	-16	4	78	78	
17	76	86	11	171	-139	5	94	76	3	217	235	8	84	91	16	59	80	5	100	-100	
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20	170	-181	14	68	77	8	427	-355	6	< 68	56	11	232	240	2	121	-122	9	98	113	
21	80	-9	15	275	-244	9	74	-59	7	< 70	-27	14	82	82	3	104	-84	10	53	-15	
22	101	89	16	< 74	-14	10	83	-74	7	85	-80	13	73	-73	4	294	246	11	-9 k 5		
0	144	-133	17	< 77	17	11	65	-70	8	80	-80	14	80	-80	5	163	135	12	159	-174	
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3	278	-278	20	109	-50	15	193	-194	12	103	-107	13	73	-73	8	58	-48	5	< 43	3	
4	117	-182	21	122	148	16	8 k 4		13	113	-104	14	44	42	10	71	64	6	56	56	
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6	313	314	1	150	-140	2	132	133	1	113	118	16	132	100	12	234	-196	8	< 47	42	
7	< 48	15	2	166	167	3	< 61	11	-12 k 4			17	145	129	13	73	51	9	< 48	38	
8	264	-267	3	189	-228	4	143	-154	1	90	75	18	145	129	14	102	102	10	50	5	
9	53	-22	4	162	-173	5	63	63	2	68	-55	19	188	153	1	39	-29	11	< 51	-45	
10	283	269	5	< 52	-9	6	94	104	3	134	165	12	204	-180	2	86	95	13	175	182	
11	97	-97	6	151	164	7	< 65	49	4	< 69	44	13	58	54	3	< 40	6	14	< 57	-39	
12	< 62	-21	7	< 56	27	8	135	-151	5	< 70	-27	14	82	82	4	< 41	9	15	< 58	22	

BASIC MERCURY(II) PERCHLORATE

Table 4. Continued.

3	<	47	26	8	68	54	0	-5	k 6	8	253	-229	1	<	51	37	3	<	24	-29	2	<	35	27														
4	<	48	39	9	57	-250	5	210	239	9	118	-127	2	<	51	-52	4	180	183	3	57	-50																
5	123	126	10	171	148	1	290	-304	10	<	48	-15	3	<	51	-49	5	150	-140	4	94	-82																
6	108	-119	11	82	62	2	37	29	11	<	50	21	4	<	52	-5	6	<	32	3	5	161	-154															
7	51	-54	12	<	45	20	3	139	130	12	77	-72	5	<	52	37	7	84	70	6	146	127																
8	82	-75	13	251	-236	4	49	-39	13	98	-99	6	101	122	8	46	35	7	100	90	7	100	90															
9	138	-160	14	129	-116	5	<	31	21	14	61	-45	7	<	58	36	9	238	205	8	79	75																
	-12	k 5	15	118	106	6	46	33	15	187	198	8	95	-91	10	122	116	9	159	171	9	159	171															
1	<	49	-35	17	190	-177	8	208	-180	9	9	k 6	9	<	60	-86	10	<	60	-86	10	<	60	-86														
2	70	68	0	5	120	-91	1	124	-126	10	124	-126	10	<	60	-86	10	<	60	-86	10	<	60	-86														
3	125	133	0	3	k 6	10	<	41	6	2	<	49	8	1	81	-80	14	<	52	-11	1	126	-101															
4	98	116	0	2	156	141	12	87	-78	3	<	50	32	2	<	22	-20	15	<	54	-42	2	82	74														
5	<	51	8	1	108	110	13	229	-220	4	<	50	15	3	<	25	-7	16	79	-42	3	<	8	7														
6	<	52	22	2	54	-61	13	229	-220	5	<	51	55	4	<	34	40	17	11	89	80	11	89	80														
7	<	53	-12	3	54	-61	13	229	-220	6	<	51	55	5	<	34	40	18	11	111	-112	12	111	-112														
8	<	54	10	4	<	34	-5	14	113	-105	7	97	-85	6	172	-177	2	52	54	4	135	-143	13	135	-143													
9	<	55	-41	5	95	97	16	74	69	8	86	86	7	85	75	3	151	-139	5	53	40	6	<	42	-30													
10	<	57	45	6	215	-230	17	131	-130	9	77	74	8	123	-128	3	54	65	6	<	42	-30	7	159	171													
11	101	-109	7	155	-171	17	131	-130	9	77	74	8	123	-128	3	54	65	6	<	42	-30	7	159	171														
	-13	k 5	8	154	158	0	6	k 6	10	9	197	202	4	54	65	6	<	42	-30	7	159	171	8	117	113													
1	75	-102	9	140	137	0	<	39	14	0	41	33	10	104	99	5	100	94	7	97	-35	8	117	113														
2	97	119	10	145	-152	1	193	-191	11	146	-86	12	<	48	2	7	<	43	21	9	<	47	34	9	126	-101												
	0	k 6	12	115	-100	3	64	62	3	<	41	25	13	<	50	44	8	55	-67	10	68	-61	10	68	-61													
0	118	142	13	99	95	4	<	41	4	4	213	209	14	69	78	10	74	80	12	<	53	20	11	96	-110													
1	<	18	3	14	9	4	143	-147	5	207	185	1	k 7	1	35	-1	12	136	127	13	87	-84	12	136	127													
2	<	33	3	15	99	-80	6	105	-100	6	212	-175	1	35	-1	12	136	127	13	87	-84	12	136	127														
3	189	-241	16	<	57	8	7	86	-86	7	92	92	2	119	132	1	4	k 7	1	59	-41	13	59	-41														
4	<	27	-27	17	8	115	124	8	121	99	3	156	-174	2	1	4	k 7	1	59	-41	13	59	-41	14	54	-44												
5	<	29	3	18	83	80	9	128	136	9	85	97	4	168	175	2	<	36	21	2	<	52	40	15	42	-82												
6	273	-281	19	<	63	-27	10	79	-70	10	178	-108	5	43	52	3	55	-49	4	87	78	16	79	-49	16	79	-49											
7	255	-230	20	117	126	11	<	50	-28	11	104	-101	2	<	51	21	3	55	-49	4	87	78	16	79	-49	17	126	-116										
8	350	297	0	-3	k 6	12	13	<	53	41	0	10	k 6	1	37	-4	4	196	-192	5	5	192	-174	17	126	-116												
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11	134	-112	2	125	-145	15	151	-154	0	<	53	52	8	<	39	6	6	50	-49	1	99	105	18	126	-116													
12	<	45	-12	3	223	-232	3	7	0	2	<	53	52	8	<	39	6	6	50	-49	1	99	105	18	126	-116												
13	153	151	4	<	57	7	0	-6	k 6	10	46	-45	9	85	82	12	<	48	22	10	72	-68	19	4	63	-35												
14	85	-69	5	101	87	0	46	-45	9	85	82	12	<	48	22	10	72	-68	19	4	63	-35	20	11	96	-110												
15	59	53	6	265	-234	2	121	-119	5	54	14	17	67	-67	11	68	55	6	<	45	-8	21	11	96	-110													
16	<	56	24	7	139	-111	3	94	-82	6	55	-53	14	14	98	12	107	113	8	<	45	-8	21	11	96	-110												
17	134	121	7	139	-111	3	94	-82	6	55	-53	14	14	98	12	107	113	8	<	45	-8	21	11	96	-110													
18	<	59	38	8	530	409	4	153	133	7	<	56	44	15	15	-22	13	5	50	-14	8	<	49	19	22	107	113											
19	<	61	-28	9	95	67	5	289	-279	8	<	57	-15	16	102	-104	14	56	45	9	<	49	19	22	107	113												
20	106	110	10	<	37	-5	6	249	-204	9	58	29	1	-1	k 7	1	156	138	4	5	k 7	1	114	-111	21	65	44											
21	65	44	11	177	-160	7	115	86	10	98	-76	1	156	138	4	5	156	138	4	5	156	138	4	5	156	138	4	5	156	138								
22	77	-115	13	177	-160	7	115	86	10	98	-76	1	156	138	4	5	156	138	4	5	156	138	4	5	156	138	4	5	156	138								
	0	k 6	14	<	30	-34	10	85	-67	1	140	143	4	175	-162	4	42	45	5	37	29	10	k 7	1	107	113	22	77	-115									
0	137	168	15	75	64	11	176	-171	2	<	44	-25	6	101	89	5	37	29	10	k 7	1	107	113	22	77	-115												
1	263	-232	16	65	14	12	156	151	3	151	-149	6	101	89	5	37	29	10	k 7	1	107	113	22	77	-115													
2	225	-210	17	65	14	12	156	151	3	151	-149	6	101	89	5	37	29	10	k 7	1	107	113	22	77	-115													
3	133	165	18	65	-65	14	12	156	151	3	151	-149	6	101	89	5	37	29	10	k 7	1	107	113	22	77	-115												
4	119	-126	19	61	-47	15	<	53	5	6	86	80	7	77	85	10	37	-44	9	<	48	-26	2	48	-26	23	117	-126	19	61	-47	15	<	53	5	6	86	80
5	98	-92	20	63	37	17	56	-27	8	7	77	85	10	37	-44	9	<	48	-26	2	48	-26	23	117	-126	19	61	-47	15	<	53	5	6	86	80			
6	33	1	21	65	37	17	56	-27	8	7	77	85	10	37	-44	9	<	48	-26	2	48	-26	23	117	-126	19	61	-47	15	<	53	5	6	86	80			
7	36	49	22	150	-190	18	102	-100	9	51	18	12	178	179	12	54	78	13	55	14	6	149	169	24	117	-126	19	61	-47	15	<	53	5	6	86	80		
8	181	171	23	150	-190	18	102	-100	9	51	18	12	178	179	12	54	78	13	55	14	6	149	169	24	117	-126	19	61	-47	15	<	53	5	6	86	80		
9	63	-58	4	169	-162	19	80	-71	12	<	55	55	15	<	54	-29	15	<	59	8	9	99	103	25	117	-126	19	61	-47	15	<	53	5	6	86	80		
10	181	175	0	273	259	0	213	204	11	12	<	55	55	15	<	54	-29	15	<	59	8	9	99	103	25	117	-126	19	61	-47	15	<	53	5	6	86	80	
11	136	132	2	103	-91	2	4	-1	19	13	64	68	15	<	56	18	16	121	-119	9	52	33	26	117	-126	19	61	-47	15	<	53	5	6	86	80			
12	104	87	3	110	106	3	134	126	14	103	103	14	103	103	14	103	103	14	103	103	14	103	103	26	117	-126	19	61	-47	15	<	53	5	6	86	80		
13	111	-115	4	217	-204	15	100	-102	16	70	-58	15	101	-116	18	67	94	1	202	243	11	<	47	-38	27	117	-126	19	61	-47	15	<	53	5	6	86	80	
14	120	-107	4	217	-204	15	100	-102	16	70	-58																											

Table 4. Continued.

0	1 k B		3 k B	2	166	163	10	< 62	0	15	< 67	100	7	262	214		
1	< 22	13	1	82	66	4	< 48	37	12	136	126	10	< 68	-69	3	104	-89
2	86	-85	2	80	-81	5	43	-45	3	58	58	11	< 64	29	16	132	85
3	178	211	3	65	-65	6	70	60	0	< 39	-75	1	< 63	-52	1	< 65	-52
4	117	138	4	89	89	7	89	94	1	133	139	2	< 63	-66	1	104	-56
5	79	-84	5	64	56	8	136	-131	2	209	207	3	< 63	-53	1	69	-55
6	< 39	1	6	< 45	7	9	< 55	-47	3	56	46	4	91	66	1	104	-56
7	81	84	7	215	-249	10	95	-50	4	180	-187	5	136	136	1	78	68
8	107	-103	8	< 49	-46	11	55	42	5	156	156	6	< 50	14	17	143	151
9	87	-83	9	124	131	12	104	-86	6	< 45	2	7	< 41	23	2	229	215
10	106	-119	10	< 53	3	13	< 63	-17	7	< 41	23	2	229	215	12	119	162
11	127	128	11	< 55	-52	14	39	32	8	< 49	54	3	93	-69	17	< 50	33
12	117	-113	12	95	96	15	< 67	-39	9	76	60	4	105	-101	10	150	-171
13	89	-105		-3 k B		16	79	-82	10	90	90	4	105	-101	10	66	70
14	79	-49		5 k B					11	56	-57	6	< 54	-35	8	< 49	29
15	71	-46	0	219	191	0	140	169	12	147	-143	7	< 55	-13	7	159	-172
16	105	105	1	133	-128	0	172	-71	13	< 60	14	8	< 57	2	8	< 48	4
17	105	-107	2	55	-45	2	70	-83	14	58	78	9	< 58	-52	9	< 48	-81
18	79	-49	3	97	111	2	74	91	15	161	-157	10	< 60	-8	10	154	-161
	-1 k B		4	71	64	3	74	91	16	149	-120	11	67	-79	11	< 48	48
0	439	-324	5	200	216	5	251	-285	16	149	-120	11	67	-79	11	137	-122
1	63	121	6	135	-121	4	72	-71	17	< 49	-26	12	155	-157	12	204	-223
2	143	221	7	167	-186	6	56	35	18	127	120	13	< 65	59	13	93	96
3	< 27	6	8	71	74	7	43	27	19	149	-120	14	< 67	31	0	93	96
4	138	-164	9	193	216	9	173	-192	0	150	-123	15	< 69	-56	1	< 47	22
5	147	-163	10	121	-103	10	79	75	1	< 56	2	16	< 70	-76	2	303	-231
6	91	89	11	< 51	-36	10	193	208	2	< 67	17	17	< 71	40	3	< 49	46
7	120	-128	12	< 54	11	12	< 55	26	3	< 57	40	18	116	129	4	153	141
8	< 42	-1	13	< 56	-4	13	< 58	4	4	< 58	17		11 k B		5	135	-111
9	56	-52	14	52	-92	14	102	-93	5	< 58	9	0	97	-84	6	110	112
10	112	117	15	110	-125	14	122	103	6	< 59	49	1	< 66	-23	7	70	-84
11	87	90	16	64	74	15	122	103	6	< 59	49	1	< 66	-23	7	68	-50
12	77	-71		4 k B					7	150	128	2	< 66	14	8	105	148
13	126	108	0	< 42	10	0	104	-85	8	< 43	52	4	< 67	-15	9	< 55	148
14	171	178	1	< 42	4	1	< 49	8	0	< 43	52	4	< 67	-15	9	105	94
15	< 60	-28	2	200	-178	2	< 50	21	1	70	-61	5	< 67	-20	3	< 55	77
16	137	-135	3	162	149	3	122	-114	2	117	-111	6	68	21	4	35	-138
	2 k B		4	114	124	4	87	77	3	103	97	7	123	80	5	35	-138
0	295	-197	5	< 46	-5	5	< 52	24	4	143	-149	8	-11 k B	8	6	106	125
1	79	65	6	< 48	27	6	< 53	-6	5	127	-131	9	< 53	50	7	106	125
2	230	196	7	106	93	7	187	-175	6	66	49	10	< 53	-41	8	119	-161
3	< 35	7	8	< 51	-36	8	< 56	-45	7	164	155	2	< 53	-22	1	232	-148
4	< 37	-34	9	< 53	-30	9	86	80	8	72	72	3	99	91	2	148	-136
5	147	-162	10	< 55	-39	10	86	-58	9	132	-137	4	126	-103	3	109	-97
6	80	67	11	76	70				10	101	167	5	< 56	-9	4	35	-138
7	39	39	12	121	126	0	315	330	11	137	133	6	< 57	35	5	109	-97
8	114	-137	13	97	-90	1	< 34	23	12	55	42	7	209	211	6	110	-96
9	67	-67	14	73	-50	2	45	-29	0	9 k B		8	< 59	-7	7	127	-103
10	< 50	-25	15	< 66	-35	3	< 37	-14	0	55	-45	9	119	-99	8	119	-99
11	92	92	16	< 68	78	4	39	44	0	66	49	10	100	94	9	88	99
12	114	-108	17	80	-79	5	95	86	2	< 60	50	11	< 64	41	10	< 54	8
13	78	72	18	92	-106	6	162	-168	3	< 60	-61	12	< 65	-13	11	< 54	-5
14	129	138		-4 k B		7	214	-219	4	61	44	13	94	-86	12	159	170
15	< 62	0	1	287	-290	8	74	57	5	79	71		-12 k B		13	63	84
16	100	-107	2	94	110	9	121	117	6	< 62	-6	0	94	96	14	127	125
	-2 k B		3	57	58	10	149	-132	7	72	-44	1	89	-88	15	127	125
0	357	300	4	192	-225	12	< 56	-32	8	65	-54	2	131	-132	16	131	120
1	< 15	-26	5	< 35	11	13	59	37		-9 k B		3	70	-58	17	< 49	2
2	< 33	-22	6	73	68	14	141	-150	0	209	227		-13 k B		18	< 49	2
3	130	170	7	42	-47	15	< 63	-11	1	80	-76	0	< 58	-39	19	< 49	-47
4	65	66	8	92	93	16	118	112	2	120	-103	1	< 59	-7	20	< 50	-9
5	216	-254	9	< 46	16				3	100	-100	2	109	105	21	< 50	-9
6	< 36	-6	10	170	181	0	< 53	-13	4	< 48	22	3	147	-143	2	< 51	-47
7	< 39	-1	11	< 52	17	1	65	-44	5	< 49	-16	4	< 61	-52	3	< 52	-44
8	90	-84	12	75	-69	2	127	-156	6	116	-118	5	< 62	25	4	< 53	19
9	166	-168	13	< 57	44	3	54	11	7	183	-181	6	< 63	-35	5	54	-55
10	81	-88	14	145	156	4	107	88	8	129	121	7	< 64	-44	6	117	-97
11	100	101	15	143	-138	5	< 55	-26	9	< 56	45	8	< 65	19	7	h 2 10	
12	< 53	-12	16	144	-150	6	< 56	18	10	< 58	-35	9	< 66	33	8	129	-163
13	< 56	-37		137	137	7	< 58	37	11	< 59	-32	10	< 67	-1	9	< 56	48
14	115	-126		137	-128	8	< 59	-5	12	< 61	32	11	< 68	-73	10	135	154
15	< 61	48	0	137	-128	9	< 60	-22	13	63	43	12	75	-72	11	105	109
16	99	102	1	< 46	34	9	< 60	-22	14	133	-128	13	81	78	12	85	-76

using anisotropic temperature factors for the mercury atoms resulted in a final R factor of 0.117 for the 2350 observed reflections. In the final cycle all shifts in parameters were less than one tenth of the calculated standard deviations.

The final parameter values are given in Table 3, and a comparison between observed and calculated structure factors is given in Table 4. Because of the presence of absorption errors in the observed F values the anisotropic temperature factors have no physical meaning and for this reason, only the isotropic temperature factors are given in Table 3.

The final parameter values were used for the calculation of a three-dimensional electron density map and a corresponding difference map. In the electron density map the peak heights were between 320 and 366 $e/\text{\AA}^3$ for the mercury atoms and 50 and 54 $e/\text{\AA}^3$ for the chlorine atoms. The average peak

height for the oxygens was $15 \text{ e}/\text{\AA}^3$ with a lowest value (for O_6) of $10 \text{ e}/\text{\AA}^3$. In the difference map all peaks, except in the immediate vicinity of the mercury positions, were below $4 \text{ e}/\text{\AA}^3$.

DISCUSSION OF THE STRUCTURE

Interatomic distances and bond angles are given in Table 5. Projections of the structure along the a and c axes are shown in Fig. 1.

Each mercury in the structure forms two short bonds to surrounding oxygens with bond lengths of 1.93 to 2.10 \AA , (average value 2.02 \AA). The bonds to the non-bridging oxygens (average bond length 1.95 \AA) are shorter,

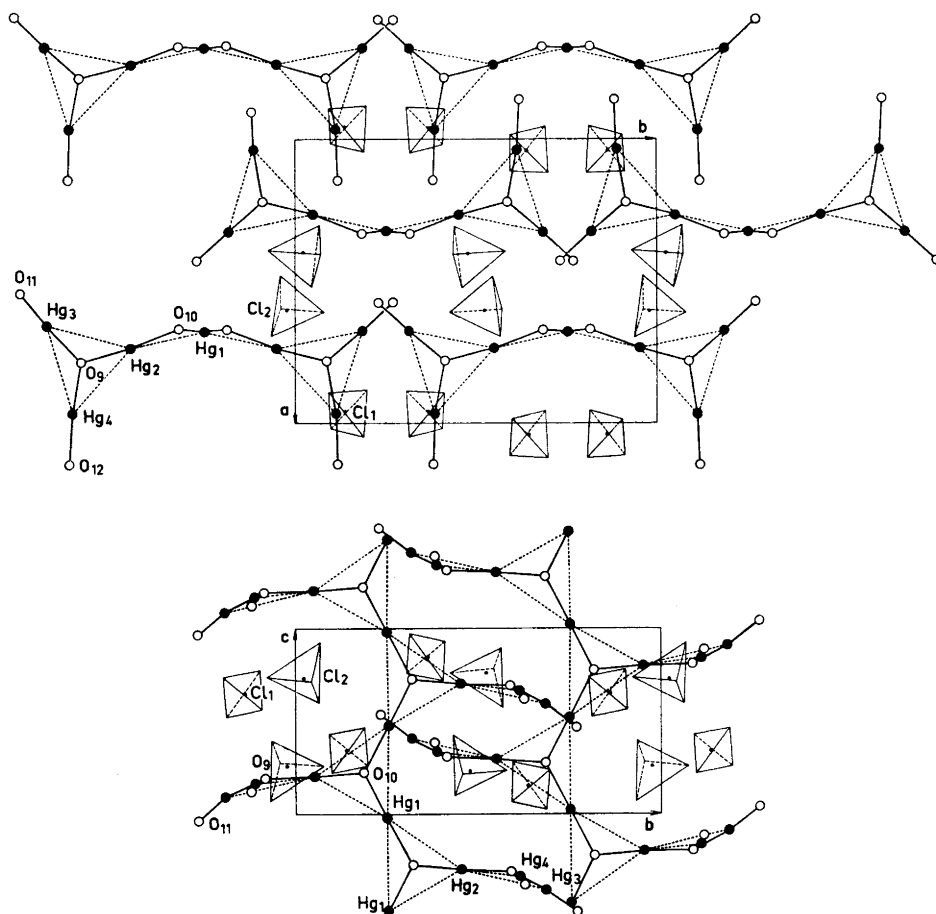


Fig. 1. Projections of the structure along the c and a axes. Full-drawn lines indicate covalent Hg—O bonds. Mercury atoms, which share covalently bonded oxygens, are connected by dotted lines.

and probably significantly so, than the bonds to the bridging oxygens (average bond length 2.04 Å). The corresponding O–Hg–O angles are close to 180° (Table 5). Other Hg–O distances are larger than about 2.7 Å and probably

Table 5. Interatomic distances and angles.

Hg–O distances (< 3.5 Å)			
Hg ₁ –O ₁₀	1.99 (3) Å	Hg ₂ –O ₉	1.98 (3) Å
–O ₁₀	2.10 (3)	–O ₁₀	2.06 (3)
–O ₁₂	2.71 (4)	–O ₁₂	2.65 (4)
–O ₈	2.78 (9)	–O ₈	2.66 (8)
–O ₁₁	2.79 (4)	–O ₆	2.70 (7)
–O ₅	2.83 (4)	–O ₁	2.84 (6)
–O ₄	3.14 (5)	–O ₅	2.85 (4)
Hg ₃ –O ₁₁	1.96 (4) Å	Hg ₄ –O ₁₂	1.93 (4) Å
–O ₉	2.05 (3)	–O ₉	2.06 (3)
–O ₁₁	2.81 (4)	–O ₁	2.72 (6)
–O ₆	2.88 (7)	–O ₂	2.91 (5)
–O ₂	3.10 (5)	–O ₃	2.92 (5)
–O ₇	3.17 (6)	–O ₁	2.99 (6)
–O ₂	3.17 (5)	–O ₄	3.13 (5)
–O ₈	3.49 (8)	–O ₄	3.14 (6)
O–Hg–O angles			
O ₁₀ –Hg ₁ –O ₁₀	171.8° (18)	O ₉ –Hg ₂ –O ₁₀	172.8° (12)
O ₉ –Hg ₃ –O ₁₁	172.9° (14)	O ₉ –Hg ₄ –O ₁₂	171.9° (15)
Hg–Hg distances (< 4.5 Å)			
Hg ₁ –Hg ₂	3.386 (2) Å	Hg ₂ –Hg ₁	3.386 (2) Å
–Hg ₂	3.507 (2)	–Hg ₄	3.398 (3)
–2Hg ₁	3.612 (3)	–Hg ₁	3.507 (2)
–Hg ₃	4.141 (3)	–Hg ₃	3.514 (3)
–Hg ₄	4.345 (3)	–Hg ₄	3.943 (5)
Hg ₃ –Hg ₂	3.514 (3) Å	Hg ₄ –Hg ₂	3.398 (3) Å
–Hg ₄	3.515 (3)	–Hg ₃	3.515 (3)
–2Hg ₃	4.070 (3)	–Hg ₄	3.942 (4)
–Hg ₁	4.141 (3)	–Hg ₂	3.943 (5)
		–Hg ₁	4.345 (3)
Hg–O distances involving O ₉ –O ₁₂ (< 4.5 Å):			
O ₉ –Hg ₂	1.98 (3) Å	O ₁₀ –Hg ₁	1.99 (3) Å
–Hg ₄	2.05 (3)	–Hg ₂	2.06 (3)
–Hg ₃	2.05 (3)	–Hg ₁	2.10 (3)
–Hg ₃	4.38 (3)	–Hg ₃	4.21 (3)
–Hg ₄	4.45 (3)		
O ₁₁ –Hg ₃	1.96 (4) Å	O ₁₂ –Hg ₄	1.93 (4) Å
–Hg ₁	2.79 (4)	–Hg ₂	2.65 (4)
–Hg ₃	2.81 (4)	–Hg ₁	2.71 (4)
–Hg ₁	4.48 (4)	–Hg ₄	4.14 (4)
		–Hg ₁	4.26 (4)

Table 5. Continued.

Hg—O—Hg angles			
Hg ₃ —O ₉ —Hg ₃	121.0° (16)	Hg ₁ —O ₁₀ —Hg ₁	123.9° (16)
Hg ₂ —O ₉ —Hg ₄	114.7 (15)	Hg ₁ —O ₁₀ —Hg ₂	113.2° (15)
Hg ₃ —O ₉ —Hg ₄	117.8° (15)	Hg ₁ —O ₁₀ —Hg ₂	114.7 (15)
Cl—O distances in the perchlorate groups			
Cl ₁ —O ₁	1.36 (6) Å	Cl ₂ —O ₅	1.42 (4) Å
—O ₂	1.29 (5)	—O ₆	1.41 (7)
—O ₃	1.27 (6)	—O ₇	1.43 (7)
—O ₄	1.41 (6)	—O ₈	1.60 (8)

correspond to non-bonding contacts. The formation of two linear covalent bonds from the mercury atom has been found in many other structures.⁸ The corresponding Hg—O bond lengths are usually found to be close to 2.0 Å.⁹

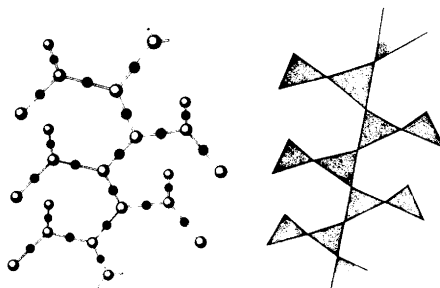
Two types of covalently bonded oxygen atoms are present in the structure. The oxygens O₉ and O₁₀ are each bound to three Hg atoms at distances between 1.98 and 2.10 Å, and O₁₁ and O₁₂ are each bound to only one Hg atom at distances of 1.96 Å and 1.93 Å. Besides being covalently bound to one Hg atom each of O₁₁ and O₁₂ is also in non-bonding contact with two other Hg atoms at distances varying between 2.65 and 2.81 Å (Table 5).

In addition to the 8 negative charges from the perchlorate groups, 24 more are needed to neutralize the positive charges of the 16 Hg atoms in the unit cell. These must be provided by the 16 covalently bound oxygen atoms (O₉—O₁₂). It seems reasonable to assume that O₉ and O₁₀, which are each bound to three Hg atoms, represent O²⁻, and O₁₁ and O₁₂, which are each bound to only one Hg atom, represent OH⁻. The correct formula for the compound should then be written Hg₂OOHClO₄.

The shortest Hg—Hg distances range from 3.39 to 3.61 Å (average value 3.50 Å) and occur between those mercury atoms, that are joined by the covalently bound oxygens O₉ and O₁₀. The next shortest Hg—Hg distances are longer than about 4.0 Å (Table 5.)

If only the mercury atoms and the covalently bound oxygen atoms are considered, the structure can be described as being built up from infinite one-dimensional complexes as illustrated in Figs. 1 and 2. Characteristic features in these complexes are the OHg₃ groups consisting of three Hg atoms at the corners of a triangle with an oxygen (O₉ or O₁₀) slightly above the center of

Fig. 2. A perspective view of one of the infinite one-dimensional complexes which can be distinguished in the structure when only the mercury atoms and the covalently bound oxygens are considered. Small filled circles indicate mercury atoms. Large circles indicate the oxygens O₉ and O₁₀, which are each bonded to three Hg atoms, or O₁₁ and O₁₂, which are each bonded to only one Hg atom. To the right, the Hg₃-triangles of the OHg₃-groups are separately drawn.



the triangle. The complexes can be described as built up from such triangles sharing corners as shown in Fig. 2. Similar OHg_3 groups can be distinguished in several other basic salts of mercury. In Hg_3OCl_4 discrete $\text{OHg}_3\text{Cl}_3^+$ ions occur¹⁰ with Hg–Hg distances of 3.53 Å and Hg–O distances of 2.05 Å, which are close to the values found here. The oxygen atom in the $\text{OHg}_3\text{Cl}_3^+$ ion is about 0.3 Å above the plane formed by the three Hg atoms.¹¹ This is found also in the present structure, where the O_9 and O_{10} atoms are about 0.3 Å (0.25 Å and 0.31 Å, respectively) above the plane of the three Hg atoms to which they are bound.

The average values for the Cl–O and O–O distances in the perchlorate groups are of the magnitude expected. The variations of the individual distances from the calculated mean are consistent with the estimated standard deviations.

As will be shown in a following paper the structure of $\text{Hg}_2\text{OOHClO}_4$ is closely related to that of the less basic $\text{Hg}_7\text{O}_4(\text{OH})_2(\text{ClO}_4)_4$. The less basic salt is found also to be built up from infinite one-dimensional complexes, which can, in fact, be described as being formed by the condensation of two of the infinite complexes in $\text{Hg}_2\text{OOHClO}_4$.

Acknowledgements. The work has been supported by *Statens Naturvetenskapliga Forskningsråd (Swedish Natural Science Research Council)*. Computer time has been made available by the *Computer Division of the National Swedish Office for Administrative Rationalisation and Economy*.

We wish to thank Mrs. Sissel Lindman for assistance in the work, and Dr. Derek Lewis for linguistic corrections.

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Received May 25, 1971.